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Effect of Salt Concentrations on the Hydraulic conductivity of the Mixtures of Basalt soil and various Bentonites

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The present paper describes the influence of the various concentrations of NaCl and CaCl₂ on hydraulic conductivity of the mixture of basalt soil and four different bentonites in the proportion of 100:20 by dry weight. The comparison of hydraulic conductivity (k) for a given salt solution on a particular soil mixture shows that the k decreased with decreasing the salt concentration. The decrease in k with a decrease in salt concentration can be attributed to increase in diffuse double thickness. The comparison of k for different salt solution shows that the divalent cation had more effect compared to the monovalent cation. However at 1 mol./L concentration almost an equal value of k was observed. With increasing the salt concentration from 0 to 0.1 mol./L, the k changed significantly which was more prominent for the sample with CaCl₂ solution. The results show that for a particular salt concentration mixture with higher liquid limit exhibited a lower k. However with increasing the salt concentration, the liquid limit converged and also the k.

INTRODUCTION

At the disposal sites of hazardous waste the concentration of the various harmful substances could exceed the concentration level that can be accepted from an environmental and health standpoint of view. The main problem is to reduce the production of dangerous leachate to such levels that the external concentration will not reach unacceptable values. This problem can be solved in several ways, one being the use of low permeable clay barriers. To prevent the ground water contamination by leachate from the waste, barriers with a low hydraulic conductivity are used as part of waste containment systems. Where suitable low hydraulic conductivity soils are not locally available, such a material can be produced by adding bentonite to a native soil. Such bentonite improved soils are used as a component of landfill liner systems.

When bentonite–soil mixtures are used for landfill liners and environmental containment barriers, they must limit the hydraulic flow. Presence of bentonite, which primarily composed of mineral montmorillonite, gives the lower value of hydraulic conductivity to the liner material. The small particle size, interlayer swelling, and thick layer of bound water associated with montmorillonite particles provides bentonite with a low hydraulic conductivity to water (Mesri and Olson 1971).

The performance of a bentonite depends upon several factors, which includes the mineralogical composition, the surface area, the surface charge deficiency of the bentonite and the composition of the exchangeable complex. In general, the performance of the bentonite increases with increase in the montmorillonite content and the surface area (i.e. decrease in the particle size). The performance also increases with an increase in the sodium (Na⁺) percentage on the exchange complex. The effect of these factors on the quality of bentonite is reflected in an increase in the cation exchange capacity, an increase in the plasticity, an increase in the swell capacity in the presence of water, and a decrease in the hydraulic conductivity when permeated with water.

However, when the bentonite comes in contact with the salt solution, these same factors make it sensitive to chemical interaction which leads to an increase hydraulic conductivity (Gleason et al. 1997; Petrov et al 1997; Ruhl and Daniel 1997).

With the growing economy of Japan the generation of municipal solid waste (MSW) has reached 50 million tones per year. Due to limited availability of space 80% of this waste has been incinerated with the residue of fly ash and bottom ash (Hanashima and Furuichi, 2000). These flyash and bottom ash are disposed in controlled landfills having a chemical compatible clay liner with a hydraulic conductivity of $<1 \times 10^{-6}$ cm/sec. Fly ash and bottom ash mostly contains cations like Na⁺ and Ca²⁺. Once it comes in contact with water, these water soluble cations moves into the clay liner and changes the properties of liner material.

For engineered landfill sites, the performance of soil liners is based on retention capacity and low hydraulic

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conductivity, which could be affected by the MSW fly ash leachate composition, in particular concentrations of cations (Mitchell 1994, Yong 1999, Warith 1987, Quigley et al 1987, 1988). Salt could affect the hydraulic conductivity of the soil liner and reduces the adsorption capacity of the heavy metals onto the soil liner (Yong and Sheremata 1991), in turn reducing the useful life of the liner. To design a secure soil liner, it is important to have a better understanding of the type and concentration of ion species that leach from the MSW ash.

Numerous attempts have been made to study the effect of these salt solutions on the bentonite alone. However in real situation a mixture of the locally available soil and bentonite is used as clay liner material. Many researchers have carried out their research work only on a particular bentonite while actually in the filed the quality and the type of bentonites vary from place to place. In addition to this the permeability test commonly used is not only time consuming but also expensive. The consolidometer permeameters system offers the means for quantitatively assessing the effect of chemical interaction on the permeability of clays (Madsen 1994). Many investigators (Newland and Alley 1960, Mesri and Olson 1971, Budhu et al 1991, Sivapullaiah et al 2000) have adopted this method, which gives quite satisfactory result of permeability as compare to the calculated values (Terzaghi 1923, Casagrande and Fadum 1944), however this method generally underestimates the value of hydraulic conductivity compared with field value (Taylor 1942, Mitchell and Madson 1987). The purpose of this investigation was to study the change in the hydraulic conductivity of mixtures of basalt and bentonites of different qualities due to permeation of salt such as NaCl and $CaCl_2$ of various concentrations using a consolidation test.

MATERIALS AND METHODS

Materials

Various mixtures of basalt soil and bentonite in a proportion of 100:20 by dry weight were used. Basalt soil was collected from Uwaba plateau of Saga prefec-

ture, Japan. Study was conducted on the mixtures of basalt soil with four different bentonites. For the sake of simplicity these bentonites named as Bentonite–A, B, C and D. The properties of all the four bentonites are tabulated in **Table 1**.

Permeant liquids

The earlier study by the author (Mishra et al, 2005) demonstrated that a concentration less than 0.1 mol/L of NaCl and $CaCl_2$ solution has insignificant effect on the behaviour of the mixtures of basalt soil and bentonite including the liquid limit, hydraulic conductivity and compressibility. Hence in the present study tests were carried out using the NaCl and $CaCl_2$ solutions of 0, 0.1, and 1 mol/L concentration on the mixtures. Solutions were prepared by dissolving salt of NaCl and $CaCl_2$ (powdered with purity grade of >95%) in deionized water. Each solution was mixed in 1 L flask. Ion chromatograph was used to measure the actual concentrations of Na^+ and Ca^{2+} present in the permeant solution. The result of the chemical analysis is tabulated in **Table 2**.

Physical and Chemical properties of mixtures

The grain size distribution of the basalt soil was obtained by the dry sieving whereas the particle size distribution was obtained by hydrometer analysis as per ASTM D 422. According to the test result 70.5% of the particles of basalt soil are less then 75 μm in diameter. The liquid limit of the mixtures was determined by both the Casagrande's method (ASTM D 4318) and falling cone method (BS 1377) by adding salt solutions with different concentration to dried soil mixtures.

Table 2. Properties of permeant liquid.

Target salt Concentration (mol/L)	Measured solute concentrations (g/L)			
	NaCl Solution		$CaCl_2$ Solution	
	Na^+	Ca^{2+}	Na^+	Ca^{2+}
1	22.7	0.9	0.006	21.1
0.1	2.2	1.3	0.002	1.6

Table 1. Properties of bentonites.

Properties	Bentonite A (Hotaka)	Bentonite B (Super clay)	Bentonite C (Kunigel VAS)	Bentonite D (Premium gel)
[+]Liquid limit (%)	310.5	615.5	391.4	511.2
[+]Plastic Limit (%)	54.1	47.6	57.2	43.9
[+]Activity	4.18	6.86	4.83	6.31
[+]Specific gravity	2.54	2.46	2.53	2.51
[+]Clay (<2μm) content	61.4	82.8	69.2	74.1
[+]Cation Exchange Capacity (cmol/kg)	52.8	72.3	61.0	92.4
*Exchangeable cation (cmolc/kg)				
Ca^{2+}	74.3	19.1	42.2	21.5
Mg^{2+}	7.5	9.6	2.2	6.7
Na^+	47.0	45.4	63.6	58.9
K^+	5.6	1.0	1.3	1.7

[+]Data provided by Hojun Co., Ltd., Japan, *Determined by CH_3COONH_4 method

The compaction curves (i.e. water content vs dry density relationship) for the mixture of basalt and bentonite was determined by adding DI water in accordance with the standard proctor test described in ASTM standard D 698.

Free swell test

The free swell test was conducted according to ASTM D 5890 for DI water and 0.1 and 1 mol./L concentration of NaCl and CaCl₂. Approximately 90 mL of DI water or salt solution was poured into a 100 mL graduated cylinder. Two grams of dry powdered bentonite was placed in the salt solution in the cylinder in 0.1 g increments. Then the cylinder was rinsed with salt solution or DI water and was filled up to the 100 mL mark. After 24 h of exposure the swollen volume of the bentonite was measured.

Consolidation test and determination of hydraulic conductivity

Consolidation tests were carried out to determine the hydraulic conductivity of the samples. The tests were carried out on the sample of 60 mm diameter and 20 mm thickness according to ASTM D 2435 using standard consolidometers. The samples were prepared by adding NaCl and CaCl₂ solutions of different concentrations to the basalt soil and bentonite mixtures, and the initial water content of the samples was adjusted to the liquid limit. The inside of the ring was smeared with a very thin layer of silicon grease in order to avoid friction between the ring and soil sample. Filter paper was placed at the bottom and top of the sample. A top cap with a porous stone was placed above the soil sample. The entire assembly was placed in the consolidation cell and positioned in the loading frame. The consolidation ring was immersed in the liquid with the same composition as the saturating fluid, and the entire consolidation cell was enclosed within a plastic bag to reduce evaporation. Then the consolidation cells were allowed to equilibrate for 24 h prior to commencing the test. All the samples were initially loaded with a stress of 4.9 kPa, increasing by an increment ratio of 1 (i.e. 4.9, 9.8, 19.6, 39.2 kPa...etc) to a maximum pressure of 1256 kPa.

Determination of hydraulic conductivity

For each pressure increment the change in the thickness of soil sample was measured from the readings of the dial gauge. Then the change in the void ratio corresponds to an increase in the overburden pressure

was calculated by the equation 1,

$$\Delta e = \frac{\Delta H}{H}$$ (Eq. 1)

Where, ΔH = Change in the thickness of sample due to increase in pressure

H = Initial thickness of the sample

From the consolidation test result, a time–settlement curve was obtained at each pressure increment. The coefficient of consolidation c_v was obtained using Taylor's square root time ($\sqrt{T}$) method.

The co-efficient of volume change can be calculated by the formula,

$$m_v = \frac{\alpha_v}{1+e}$$ (Eq. 2)

where α_v = coefficient of compressibility

$$= \frac{\Delta e}{\Delta \sigma} \quad \text{where,}$$

$\Delta \sigma$ = Change in pressure
Δe = Change in void ratio

The coefficient of permeability, k, was calculated using the equation–3 for various pressure increments using the c_v, and coefficient of volume change, m_v,

$$k = c_v \, m_v \, \gamma_w,$$ (Eq. 3)

where, γ_w is the unit weight of the pore fluid

RESULTS AND DISCUSSION

Liquid Limit

The effect of the pore water on the liquid limit of the mixtures is shown in **Table 3**. The liquid limit of the mixtures is seen to decrease with increasing salt concentration. For a given concentration, the NaCl salt solution gives a higher liquid limit than the CaCl₂ solution. This trend was quite consistent with diffuse double layer theory. Interparticle forces have a more prominent role in determining the liquid limit. The distance between particles or between structural units of the particles is such that the forces of interaction between the clay particles become sufficiently weak to allow easy movement of particle relative to each other. The soils are water saturated at this point, and the cohesion between particles in remoulded sample is small. In high swelling clays such as montmorillonite, the dominant interparticle force is one of repulsion. This force of repulsion determines the distance between particles. Therefore

Table 3. Liquid limit for the different salt concentrations for the mixtures of basalt soil and various bentonites.

Salt concentration (mol./L)	Mixture of basalt with different bentonites (in 100:20 proportion)							
	with bentonite–A		with bentonite–B		with bentonite–C		with bentonite–D	
	NaCl	CaCl₂	NaCl	CaCl₂	NaCl	CaCl₂	NaCl	CaCl₂
0 (DW)	78.4	–	140.5	–	92.7	–	118	–
0.1	68.5	67	103.8	81.8	72.4	67.5	88.5	69.2
1	63.3	63	73.5	73.3	62.7	61.3	62.3	61.3

an increase in the salt concentration or substitution of divalent for monovalent exchangeable cation decreases the inter–particle repulsion. As a result the particle becomes free to move at lower water contents or lower inter–particle distances, and the liquid limit decreases (Warkentin, 1961).

Table 3 also shows that the decrease in the liquid limit due to the increase in the salt concentration is small for a range of 0.1 to 1 mol/L than for the range of 0 to 0.1 mol/L. An increase in the CaCl$_2$ concentration from 0 to 0.1 mol/L produced a large decrease in the liquid limit compared to NaCl solution. A further increase in the CaCl$_2$ concentration did not produce any significant decrease in the liquid limit.

A comparison of the liquid limit of the mixtures show that, the mixture with bentonite–B possess the highest liquid limit which can be attributed to the clay fraction percentage, i.e. a large surface area (Bojana et al 2004, Muhantham 1991). The change in the liquid limit due to the salt concentration was higher for the bentonites with higher initial liquid limit before addition of salt. The mixture with Bentonite–B exhibited a highest reduction in the liquid limit by adding salt solution.

Free swell

The effect of NaCl and CaCl$_2$ concentrations on free swelling for the four pure bentonites is shown in Table 4. At the equal concentration higher swelling was observed for NaCl compared to CaCl$_2$ solutions. For NaCl solutions osmotic as well as hydration swelling takes place which allows the interlayer spacing to become large, resulting in greater swelling. On the other hand only hydration swelling takes place for CaCl$_2$ solutions (Norrish and Quirk, 1954; Zhang et al, 1995). With increasing the salt concentration, the swell volume decreases. When the cation concentration in the bulk solution increases, water leave the interlayer region due to the gradient of free energy induced by the elevated concentration in the bulk pore water.

A significant reduction in swelling took place when the salt concentration was increased from 0.1 to 1 mol/L. At 1 mol/L of the NaCl and CaCl$_2$ solution nearly equal swelling was observed for all bentonites. The decrease in the swelling by the increase of the concentration from 0 to 0.1 mol/L was higher for CaCl$_2$ than for NaCl.

Comparison of the swelling for the bentonites show that bentonite–A with the highest amount of exchange-

able calcium exhibits the least swelling in water. The bentonite–B and D with an equal exchangeable sodium percentage (ESP) (i.e. 60 and 66% respectively) exhibits the same amount of swelling. For bentonite–C with an ESP of 58% less swelling occurred compared to bentonite B and D. The less swelling of bentonite–C can be attributed to presence of a higher amount of exchangeable calcium. Data from the table also shows that there is a significant reduction in the swelling occurred when the CaCl$_2$ concentration increased from 0.01 to 0.1 mol/L. On the other hand the swelling decreased significantly with increasing the NaCl concentration from 0.1 to 1 mol/L. These two findings show that a higher salt concentration of NaCl is required to replace the exchangeable cations as compare to CaCl$_2$ solution. This is due to the higher valency of calcium which replaces the monovalent exchangeable cation easily compared to sodium ion.

Hydraulic conductivity

Hydraulic conductivity of the mixtures was calculated for various pressure increments using experimentally determined values of c_v and m_v. Figure 1 show the relationship between the void ratio and hydraulic conductivity (k) at different NaCl and CaCl$_2$ concentrations for the four different mixtures. It can be seen that log k varied approximately linearly with the void ratio. Similar observation has been reported by other investigators (Olson and Daniel, 1981; Pandian et al, 1995) for the samples permeated with pure water. Each of the plots shows that the k decreases with decreasing void ratio. Anisotropic consolidation of the samples causes reorientation of the clay platelets into a plane perpendicular to the direction of the major principle stress and there by increases the tortuosity factor (Quigley et al, 1966) which results in the reduction in k. The compression of the soil sample in the consolidation cell due to increase in the overburden pressure resembles the compression of the liner material caused by increase in the overburdened waste. Hence the hydraulic conductivity of the liner decreases by an increase of the tortuosity factor induced by the increase of the overburden waste.

Comparison of the figures shows that Ca–soil exhibits a higher k values than Na–soil at a given concentration for a particular sample. However the k was almost identical at 1 mol/L concentration for both NaCl and CaCl$_2$ solution. Jo et al. (2001) had also reported similar observation. Similar observation for the liquid

Table 4. Free swelling for the different bentonites at various salt concentrations.

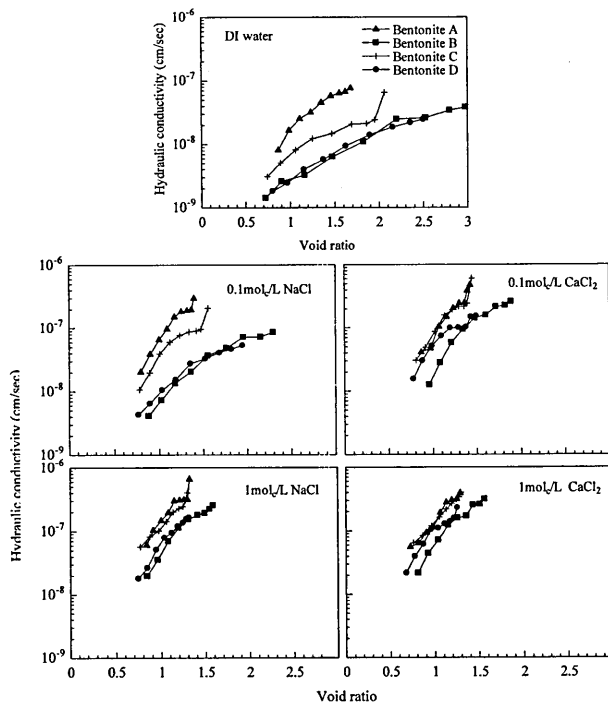| Salt concentration (mol/L)	Free swell for pure bentonite (mL/2 g)							
	Bentonite–A		Bentonite–B		Bentonite–C		Bentonite–D	
	NaCl	CaCl$_2$	NaCl	CaCl$_2$	NaCl	CaCl$_2$	NaCl	CaCl$_2$
0 (DW)	14	–	25	–	21	–	24	–
0.01	13	12	22	21	20	18	22	19
0.1	12	9	20	13	16	10	20	11
1	7	7	8	8	6	6	8	7

Fig. 1. Plots for the hydraulic conductivity versus void ratio at various concentration of NaCl and $CaCl_2$ for the mixtures of basalt and various bentonites.

Fig. 2. Plot for hyadraulic conductivity and void ratio for the mixture of basalt and different betnonites for different concentration of NaCl at void ratio corresponds to the maximum dry density.

Fig. 3. Plot for hyadraulic conductivity and void ratio for the mixture of basalt and different betnonites for different concentration of $CaCl_2$ at void ratio corresponds to the maximum dry density.

limit as well as the free swell at 1 mol./L concentration was also observed. All figures indicate that the k changed significantly when the concentration increased from 0 to 0.1 mol./L and this change was more pronounced for $CaCl_2$ than for NaCl. The results for all the mixtures show that increase of the concentration from 0 to 0.1 mol./L increased the k by 2 to 4 times for NaCl solution, while increased in k for $CaCl_2$ by 6 to 14 times. However with further increasing the concentration from 0.1 to 1 mol./L, k increased by only 1.1 to 1.7 times for $CaCl_2$ but 3 to 5 times for NaCl. Similar behavior was also observed for the liquid limit as well as for the free swelling of the bentonites.

Comparison of the different bentonites for a particular salt concentration shows the sample with higher liquid limit has a lower k. The mixture with bentonite–B which has the highest liquid limit exhibited lowest k while for the mixtures with bentonite–A, which has a lower liquid limit exhibited highest k for all the salt concentration range. However with increasing the salt concentration, the liquid limit for all the mixtures converges and so the k. It is also observed that for same concentration the difference in k for all the mixture for $CaCl_2$ is less compared with NaCl.

Comparison of k for different bentonites at a given void ratio

Figures 2 & 3 show the salt concentration versus hydraulic conductivity at the void corresponding to the maximum dry density and optimum moisture content prepared from **Figure 1**. Figures shows that increase in the salt concentration enhance the k of the samples. This increase was quite prominent for $CaCl_2$ compared
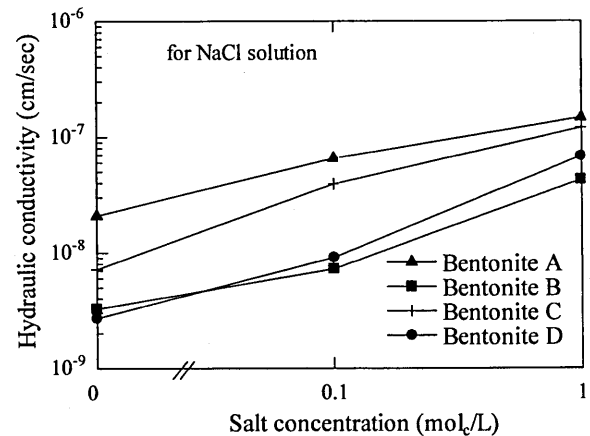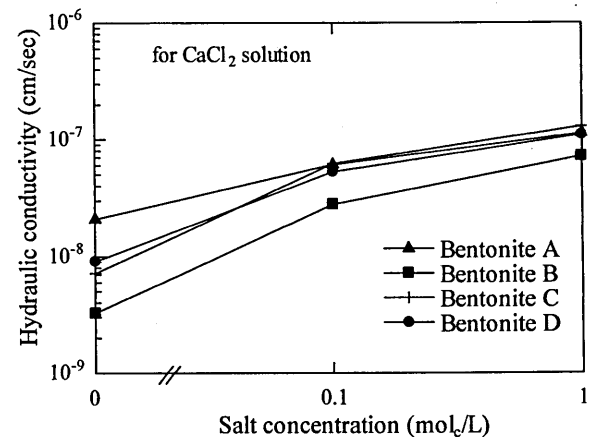
to NaCl solution. Comparison for the different concentration of NaCl shows that k for the mixture containing bentonite–D increased to a large extent as comparison to other. The k increased 25 times when the NaCl concentration increased from 0 to 1 mol./L for the mixture with bentonite–D. However the corresponding increase was only 7 times for the mixture with bentonite–A. Similar comparison for the different concentration of $CaCl_2$ shows that an increase in the concentration from 0 to 1 mol./L enhanced the k for the mixture with bentonite D by 40 times, which was the highest among the mixtures. Similarly mixture with bentonite–A showed a marginal increase in the k. The higher change in the k for the bentonite–D is due to its higher ESP and cation exchange capacity.

SUMMARY AND CONCLUSION

Consolidation test was performed on the various mixtures of basalt soil and bentonite in the proportion of 100:20 by dry weight to evaluate the effect of salt on the

hydraulic conductivity of the mixtures. The results demonstrated that the salt concentration has a pronounced effect on the hydraulic conductivity of the mixtures.

When the salt concentration was increased from 0 to 0.1 mol,/L, the change in the k was significant which was more pronounced in CaCl$_2$ than NaCl. An increase of the concentration from 0 to 0.1 mol,/L raised k by 2 to 4 times for NaCl solution whiles 6 to 14 times for CaCl$_2$ in all the mixtures. For a particular salt concentration the mixture with a higher liquid limit exhibited a lower k. However with increasing the salt concentration the difference between the liquid limit as well as k for all four mixtures reduced. When the k for the mixtures with different bentonites is compared at the same salt concentration the mixture containing the bentonite with a higher liquid limit showed a lower k. There is less difference in the k of the mixtures for the higher salt concentration. The difference in the k of the mixtures at the same concentration is smaller for CaCl$_2$ than for NaCl.

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